**p-Xylylenediamine and Its New Polyimides**

Bulat A. Zhubanov¹, Valentina D. Kravtsova*, Kair A. Zhubanov², 
Tleutay S. Abildin² and Nagima B. Bizhanova²

¹Institute of Chemical Sciences, 106 Walikhanov St., 480100 Almaty, Kazakhstan 
²Scientific Institute of New Chemical Technologies and Materials 
95a Karasay Batyr, 480012 Almaty, Kazakhstan

**Abstract**

*p*-Xylylenediamine was synthesized with high yield through a hydrogenization of dinitrile of terephthalic acid. The effect of catalyst and solvent on the yield of the product was studied. The hydrogenization was carried out on skeletal catalysts based on alloy of Reney nickel Ni:Al (1:1) in aliphatic alcohols (C₁-C₄) at 4.0 MPa and 60°C. New alkanearomatic polyimides were synthesized by one-step polycondensation of *p*-xylylenediamine and dianhydride of tricyclodecendecarboxylic acid in the presence of isonicotinic acid as a catalyst in nonpolar amide solvents at 110-120°C. The films based on the synthesized polyanilimides are stable with 70-80 MPa breaking strength and 30-40% elongation. The glass transition temperature of the polymers is within 265-280°C temperature range, the temperature of decomposition is above 340°C. The tangent of dielectric loss of the films at 1 kHz and 25°C is 0.002-0.004, the dielectric permittivity is about 3.22-3.35.

**Introduction**

Polyalkanimides (PAIs) is a promising material for electro- and radiotechnic utilization [1,2]. Being between aromatic and aliphatic polymers, PAIs possess properties of both type of polymers, having aromatic polyheterocycles with aliphatic groups. The introduction of a methylene chain to a polymer leads to the decrease of its melting point but promotes an increase in hydrolysis stability. That insures easy production such polymers in industry [2].

There is a number of PAIs on the basis of pyromelkyte dianhydride and aliphatic diamines – deca-, dodecamethylenediamines and their mixtures. They manifest good performances during long period of their exploitation with high temperature above 200°C [2,3]. However, we consider in this paper the synthesis of PAI based *p*-xylylenediamines (*p*-XDA) with various dianhydrides.

Earlier we have reported the development of the new monomer – dianhydride tricyclodecendecarboxylic acid at the Institute of Chemical Sciences (Scheme 1). On the Scheme 1 R: H (Adduct of benzene - AB), F (Adduct of fluorobenzene - AFB), Cl (Adduct of chlorobenzene - ACB), which are initial com-

**Sheme 1**

Institute of New Chemical Technologies and Materials developed the new procedure for the preparation of *p*-XDA with high yield on Reney nickel catalysts. This paper describing synthesis of polyimides on its basis and some physical properties of materials from the polyimide.

**Experimental**

Purification of original agents and solvents was carried out according to the standard procedures. IR-spectra were recorded on Jasco IR-810 as KBr disks for monomers and films for polymers.

Viscosity of 0.5 wt.% polymer solutions was measured with Ubellode capillary viscosimeter at 25°C in DMSO.

Thermogravimetric analysis was performed with
TGA Metler Toledo at 8°C per min and DSC Metler Toledo for $T_g$ analysis. The loss of weight was calculated from TGA data at loss of 0, 5, 10 and 50% of weight ($T_0$, $T_s$, $T_{10}$, $T_{50}$).

Breaking strength $\sigma$, relative elongation $l$, tangent of dielectric loss angle $\tan \delta$, and dielectric permeability $\varepsilon$ were measured according standard procedures using GP-3 machine [6,7].

$p$-XDA was synthesized by catalytic reduction of $p$-dinitrile of terephthalic acid ($p$-DTA)\(^5\). The hydrogenization was carried out in the kinetic reactor under high pressure [8] with control of hydrogen supply [9]. An alloy of Rene nickel based on Ni:Al (1:1) was used as a catalyst. The reactor was a cylinder with volume of 0.15 L under shaking with frequency of 600 per min. 50 mL methanol as a solvent saturated by ammonia at cold conditions. The catalyst (0.5 g) was immersed in the reactor under the solvent with following gradual injection of 1.43 g of $p$-DTA. The hydrogenization was carried out until finishing of its absorption from gas phase by the solvent.

The composition and the structure of $p$-XDA were identified by element analysis and IR-spectroscopy. Measured: C\(_8\)H\(_{12}\)N\(_2\), %: C 70.26; H 8.84; N 19.92. Calculated: %: C 70.60; H 8.09; N 20.57.

IR-spectra cm\(^{-1}\): 1500 (benzene ring), 1670, 3420 - NH\(_2\)-group, 2920 - CH in -CH\(_2\)-group.

Melting point is 34.5-35°C (ref. 35°C [8]).

Dianhydride of tricyclo-[4,2,2,0\]2,5-dec-7-en-3,4,9,10-tetracarboxylic acid (photoadduct of benzene and maleic anhydride - AB), and dianhydrides of 7-fluorotricyclo-[4,2,2,0\]2,5-dec-7-en-3,4,9,10-tetracarboxylic acid (AFB) and 7-chlorotricyclo-[4,2,2,0\]2,5-dec-7-en-3,4,9,10-tetracarboxylic acid (ACB), synthesized according to our procedure [10,11], were purified 3-4 times washing in hot acetone during 2 hrs. The precipitated monomers were filtered followed drying at 80-90°C.

AB m.p. – 250-252, AFB m.p. – 312-314, and ACB m.p. – 302-304°C.

Composition of monomers is coincided with theoretical calculation. Polyimide films were prepared by casting of their 25 wt.% DMAA solution on a glass followed drying at 80°C within 20 min and final drying during 1 hr at 150°C.

Polyimide based on $p$-XDA and ACB was prepared with a yield 98.5-99.0% and degree of imidization about 100%.

**Results and Discussion**

The catalytic hydrogenization of nitriles is a complicated process with a number of parallel-sequent reactions [8,13,14]. The reduction of $p$-DTA was carried out in the presence of catalysts, which are summarized in Table 1.

### Table 1

<table>
<thead>
<tr>
<th>No</th>
<th>Composition of initial alloy before leaching (wt.%)</th>
<th>Duration (min)</th>
<th>$p$-XDA yield (wt.%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>Ni:Al = 50.0:50.0</td>
<td>140</td>
<td>73-75</td>
</tr>
<tr>
<td>2</td>
<td>Ni:Ti:Al - industrial</td>
<td>50</td>
<td>90-92</td>
</tr>
<tr>
<td></td>
<td>Al-50:53.5; Ni-44:46.5; Ti-2.2-2.8</td>
<td></td>
<td></td>
</tr>
<tr>
<td></td>
<td>H-5 catalyst</td>
<td></td>
<td></td>
</tr>
<tr>
<td>3</td>
<td>Ni:Me:Al = 47.5:2.5:50.0</td>
<td>41</td>
<td>94-95</td>
</tr>
<tr>
<td>4</td>
<td>Ni:Me:Al = 45.0:5.0:50.0</td>
<td>32</td>
<td>97-98</td>
</tr>
</tbody>
</table>

Hydrogen pressure – 4 MPa, temperature 60°C

The data of the hydrogenization of $p$-DTA to $p$-XDA was shown as a reference in Table 1 and 2. Table 1 proves that catalysts Ni-Ti and H-5 are the most active and selective in comparison with Ni-skeletal one. The yield of the product on Ni-Ti is 90-92% and on H-5 is about 97-98%. From this point we carried out the synthesis with H-5 catalyst. It is known that the hydrogenization of aromatic dinitriles is mostly affected by the type of solvent and the best
results were achieved in the presence of methanol [14].

In this paper we studied especially the effect of solvent on the hydrogenization of \( p \)-DTA on H-5 catalyst in the presence of ammonia. The data are listed in Table 2. It is known that the maximal yield is reached at stoichiometric ratio of components on a catalyst surface [15]. That was taken into account during the synthesis.

The highest yield was obtained at ratio of nitrile: ammonia = 1:3 in alcohol solution.

The presence of aminonitriles among the products of nitrile hydrogenization points out that sequent reduction of nitrile groups occurs [14,16,17]. There is a reducing rate of \( p \)-DTA hydrogenization on H-5 catalyst accompanied with sorption of the calculated amount of hydrogen [18]. The hydrogenization at appropriate speed occurs until beginning sorption of 2 mol hydrogen per 1 mol of dinitrile. Then the rate decreases and the following hydrogen reacts at lower speed [17,18].

The hydrogenization of \( p \)-DTA could be schematically described as a hydrogenization of a nitrile group of dinitrile with the formation of aminonitrile (cyanbenzylamine). Then it forms a final \( p \)-XDA after hydrogenization of the second nitrile group.

The increase of catalyst activity could be achieved by changing of by different ratios of nickel and alumina components from NiAl\(_3\) and Ni\(_2\)Al\(_3\) alloys to NiAl\(_3\) one [19]. Thus, the product yield and the reaction rate are strongly dependent on the type of catalyst and synthesis conditions. High activity and selectivity of H-5 catalyst are due to high degree of its hydrogen sorption, which is 2-3 times higher then that of Ni one [19].

Dianhydrides of tricyclodecetetracarboxylic acid are the initial monomers for the preparation of various thermostable polymers, including polyimides [20,21]. Earlier polyimides based on \( p \)-XDA and the dianhydride could not be synthesized efficiently due to the absence of any effective synthetic procedures (Scheme 2).

In this work the polyimides were polymerized by one-step polycondensation in polar non-proton solvent of amide in the presence of isonicotinic acid (pyridine-4-carboxylic acid) as a catalyst. This agent as any pyridinecarboxylic acid speeds up essentially the polycondensation [5]. It is known that the polyheterocycles prepared in liquid medium show an appropriate solubility in comparison with polyamide acid) synthesized in solid phase. Among DMFA, DMSO, MP and DMAA solvents, the last one manifests the highest viscosity of synthesized polyimide solution. For instance, the viscosity of 0.5 wt.% solutions of fluoro-containing polyimide at optimal conditions is: DMFA – 0.45; DMSO – 0.75; MP – 0.53; and DMAA – 1.23 dL per g. The further investigations were carried out using DMAA solvent.

The study of interaction between \( p \)-XDA and alicyclic dianhydrides shows that the most active dianhydride is AFB. That is due to the presence of fluorene atom at the endoethylene bond. Fluorine having the highest electonegativity is a reason of the increase of electrophilicity of the anhydride groups. The activity is typical for AB and ACB [10]. The polymerization was carried out at 110-120°C within 2-2.5 hrs with 100% imidization. The viscosity of resulting PAIs is 0.94-1.23 dL per g range. The optimal conditions of PAI synthesis are summarized in Table 3 in comparison with the data on two-step synthesized polyimide based on AFB and \( p \)-XDA (number 4). The one-step polyimides show higher performances of their viscosity then two-step one.

The composition and the structure of PAIs were identified by elemental analysis and IR-spectrometry. There are bands of polyimides in IR-data at 1775

### Table 2

<table>
<thead>
<tr>
<th>No</th>
<th>Solvent: alcohol + ammonia</th>
<th>Catalyst Ni-Ti (= 3% Ti)</th>
<th>Catalyst H-5</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Duration (min) p-XDA yield, (wt.%)</td>
<td>Duration (min) p-XDA yield, (wt.%)</td>
<td></td>
</tr>
<tr>
<td>1</td>
<td>methanol 50 90-92</td>
<td>32 97-98</td>
<td></td>
</tr>
<tr>
<td>2</td>
<td>ethanol 62 89-91</td>
<td>41 96-97</td>
<td></td>
</tr>
<tr>
<td>3</td>
<td>propanol 69 90-91</td>
<td>48 96-97</td>
<td></td>
</tr>
<tr>
<td>4</td>
<td>n-butanol 71 90-91</td>
<td>53 96-97</td>
<td></td>
</tr>
</tbody>
</table>

Hydrogen pressure – 4 MPa, temperature 60°C.
Scheme 2

Table 3

<table>
<thead>
<tr>
<th>No</th>
<th>Dianhydride</th>
<th>C&lt;sub&gt;res&lt;/sub&gt;/C&lt;sub&gt;catalyst&lt;/sub&gt;</th>
<th>T, °C</th>
<th>Time, hr</th>
<th>η (0.5%, DMSO, 25°C), dL/g</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>AFB</td>
<td>40/2.0</td>
<td>120</td>
<td>2.2</td>
<td>1.23</td>
</tr>
<tr>
<td>2</td>
<td>AB</td>
<td>40/2.5</td>
<td>120</td>
<td>2.5</td>
<td>1.10</td>
</tr>
<tr>
<td>3</td>
<td>ACB</td>
<td>40/2.5</td>
<td>110</td>
<td>2.5</td>
<td>0.94</td>
</tr>
<tr>
<td>4</td>
<td>AFB 2-step</td>
<td>20/0</td>
<td>25</td>
<td>2.5</td>
<td>0.32</td>
</tr>
</tbody>
</table>

and 1715 (carbonyl of imide cycle), 1365-1370 (tret-N) and 715-725 cm<sup>-1</sup> (imide cycle). Fluoro-containing PAI shows C-F peak at 1260-1300 and chloroone – C-Cl at 860 cm<sup>-1</sup>. The absorbance bands of

non-cycled amide acids are not present.

The presence of alicyclic structure as well as fluo-
rine and chlorine atoms in the case of AFB and ACB
gives rise to their decreased the glass transition point.
The presence of halogen atom decreases the glass
transition point of PAI by 10-15°C approximately,
besides a decrease by chlorine is more significant
than that by fluorine presence (Table 4).

| Poly-
mer | $T_p$°C | Temperature of weight loss, °C | air | argon |
<table>
<thead>
<tr>
<th></th>
<th></th>
<th></th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>$T_0$</td>
<td>$T_5$</td>
<td>$T_{10}$</td>
<td>$T_{50}$</td>
</tr>
<tr>
<td>AB</td>
<td>270</td>
<td>340</td>
<td>380</td>
<td>420</td>
</tr>
<tr>
<td>ACB</td>
<td>280</td>
<td>325</td>
<td>350</td>
<td>400</td>
</tr>
<tr>
<td>AFB</td>
<td>265</td>
<td>335</td>
<td>365</td>
<td>415</td>
</tr>
</tbody>
</table>

The temperature of PAI decomposition in air is
the narrow range within 325-340°C, however halogen-
containing polyimides decompose at lower tem-
perature. The thermal stability of PAI in argon is ob-
viously higher. The lag between $T_p$ and $T_g$ is about
45-70°C that points out the possibility of their pro-
cessing under pressure.

The synthesized polymers demonstrate a good vis-
cous performance, and an appropriate solubility. That
makes able to cast excellent films with breaking
strength 70-80 MPa and elongation 30-40%. Module
of elasticity in the dependence of modifiers (triphenylphosphate, dimethyl- and dibutylphthalate,
dimethylterephthalate et al.) is limited within 2700-
3400 MPa

The tangent of dielectric loss angle at 1 kHz, 25°C
and zero humidity is 0.002-0.004, and dielectric
permittivity is about 3.22-3.35. Low performance of
dielectric permittivity is due to the effect of both ali-
cyclic unit and electronegative halogen atom. That
shows the better results in the case of aromatic poly-
imides.

Conclusions

$p$-XDA with high yield was synthesized by hydro-
genization of $p$-DTA in the presence of skeletal cata-
lyst based on Reney nickel from Ni:Al alloys in ali-
phatic alcohol solvents. A series of new non- and
halogen-containing PAI were polymerized by poly-
condensation of $p$-XDA with various dianhydrides
of tricyclodecetetracarboxylic acid in polar non-pro-	on solvents of amide type at the presence of isonico-
tinic acid as a catalyst. These polymers demonstrate
good physico-mechanical and dielectric performan-
ces as well as a high thermal stability. Due to the
high content of functional groups these PAIs could
be a good candidate for the preparation of metal-cont-
taining polyimide films for application in electron-
ics and aerospace industry.

Acknowledgment

$p$-DTA was obtained as a kind gift from Dr. P. B.
Vorobiev and engineer L. F. Gabdullina (Institute of
Chemical Sciences).

References

1. Kazaryan L. G., Azriel A.E., Vasileva V.A. et
2. Kalugina E.V., Blumenfeld A.B., Novotortsev
3. Chernova A.G., Savina M.E., Pinaeva N.K. et
A. New thermostable heterocycle polymers.
Alma-Ata, Nauka, 1979, p. 252.
6. Kalinina L.S., Motorina M.A., Nikitina N.I.,
Khachapuridze N.A. Analysis of condense poly-
7. Bradulina L.G., Gavrilova N.D., Vygodskiy Y.S.,
Matieva A.M. Vysokomol. Soed. 41B(5):901
(1999).
et al. Izvestia ANKazSSR, Ser. Chem. 4:68
10. Zhubanov B.A. Polymer Yearbook, v.4, 1987,
p. 149.
11. Zhubanov B.A., Kravtsova V.D. Izvestia NAN
12. Nikolskiy B.P. (ed.) Spravochnik Chimica, V.2,
15. Sokolskiy D.V. Hydrogenization in solutions.


Received 28 June 2003.